



(12) **EUROPEAN PATENT APPLICATION**

(43) Date of publication:  
25.08.2004 Bulletin 2004/35

(51) Int Cl.7: **C07C 253/30**, C07C 209/48,  
C07C 211/27, C07C 211/30

(21) Application number: **04002100.8**

(22) Date of filing: **31.01.2004**

(84) Designated Contracting States:  
**AT BE BG CH CY CZ DE DK EE ES FI FR GB GR**  
**HU IE IT LI LU MC NL PT RO SE SI SK TR**  
Designated Extension States:  
**AL LT LV MK**

(30) Priority: **20.02.2003 JP 2003042397**

(71) Applicant: **mitsubishi gas chemical**  
**COMPANY, INC.**  
**Chiyoda-ku, Tokyo (JP)**

(72) Inventors:  
• **Kanamori, Yoshinori,**  
**Mitsubishi Gas Chem. Co., Inc.**  
**Niigata-shi Niigata (JP)**

• **Ebata, Shuji, Mitsubishi Gas Chem. Co., Inc.**  
**Niigata-shi Niigata (JP)**  
• **Tsukahara, Kengo,**  
**Mitsubishi Gas Chem. Co., Inc.**  
**Niigata-shi Niigata (JP)**  
• **Yamamoto, Yoshiaki,**  
**Mitsubishi Gas Chem. Co., Inc.**  
**Niigata-shi Niigata (JP)**

(74) Representative:  
**Gille Hrabal Struck Neldlein Prop Roos**  
**Patentanwälte,**  
**Brucknerstrasse 20**  
**40593 Düsseldorf (DE)**

(54) **High-selective production method of di(aminomethyl)-substituted aromatic compound**

(57) A di(aminomethyl)-substituted aromatic compound is produced by a two-stage hydrogenation. In the first stage, an aromatic dinitrile is hydrogenated into a cyano(aminomethyl)-substituted aromatic compound in the presence of a Pd-containing catalyst. The resultant cyano(aminomethyl)-substituted aromatic compound is then hydrogenated in the second stage into the target

di(aminomethyl)-substituted aromatic compound in the presence of a Ni- and/or Co-containing catalyst. By the above method, the di(aminomethyl)-substituted aromatic compound is produced in a high selectivity and a sufficiently high yield without reducing the catalyst life.

## Description

### BACKGROUND OF THE INVENTION

#### 1. Field of the Invention

[0001] The present invention relates to a production method of a di(aminomethyl)-substituted aromatic compound by the hydrogenation of an aromatic dinitrile. The di(aminomethyl)-substituted aromatic compound is useful as the raw material for producing hardeners, synthetic resins, isocyanates, etc.

#### 2. Description of the Prior Art

[0002] It has been proposed to hydrogenate an aromatic dinitrile in the presence of catalyst system containing various metals. For example, Japanese Patent Publication No. 51-24494 proposes a production method of 3-cyanobenzylamine or 4-cyanobenzylamine by the palladium-catalyzed hydrogenation of isophthalonitrile or terephthalonitrile in the presence of liquid ammonia and an inorganic alkali. Japanese Patent Application Laid-Open Nos. 9-40630 and 10-204048 propose production methods of aromatic cyanomethylamine (cyano(aminomethyl)-substituted aromatic compound) by hydrogenating one of the nitrile groups in an aromatic dinitrile in the presence of a nickel and/or cobalt-containing Raney catalyst. These patent documents only relate to the production of the aromatic cyanomethylamine, and describe nothing about the production of a diamine (di(aminomethyl)-substituted aromatic compound) corresponding thereto.

[0003] Japanese Patent Publication No. 10-502671 discloses a semi-hydrogenation of an aliphatic dinitrile into a corresponding aliphatic aminonitrile in the presence of a catalyst selected from the group consisting of Raney nickel catalysts containing a doping element such as zinc and elements in groups IVb, VIb, VIIb and VIII of the periodic table and Raney cobalt catalysts containing a doping element such as zinc and elements in groups IVb, VIb, VIIb and VIII of the periodic table. Japanese Patent Application Laid-Open No. 2001-524464 proposes a continuous partial hydrogenation of an aliphatic dinitrile into a corresponding aliphatic aminonitrile in a heterogeneous system in the presence of a hydrogenation catalyst. These patent documents only relate to the partial hydrogenation of the aliphatic dinitrile into the aliphatic aminonitrile, and describe nothing about the partial hydrogenation of the aromatic dinitrile into the aromatic cyanomethylamine and the subsequent hydrogenation of the aromatic cyanomethylamine into the corresponding diamine (di(aminomethyl)-substituted aromatic compound).

[0004] It has been further proposed to convert the aromatic dinitrile to the corresponding diamine (di(aminomethyl)-substituted aromatic compound) by a single step hydrogenation in the presence of a Ni- or Co-con-

taining catalyst. For example, Japanese Patent Publication No. 38-8719 proposes the hydrogenation of phthalonitrile into the target xylylene diamine in alcohol containing a slight amount of water in the presence of a Raney nickel or Raney cobalt catalyst doped with caustic alkali. Japanese Patent Publication No. 48-22593 proposes the production of xylylene diamine from phthalonitrile in the presence of a supported nickel catalyst containing magnesium oxide as the co-catalyst. However, these methods fail to produce the target diamine in a sufficiently high yield because of the formation of by-products.

### SUMMARY OF THE INVENTION

[0005] An object of the present invention is to provide a novel method for producing the di(aminomethyl)-substituted aromatic compound in a high selectivity.

[0006] As a result of extensive study in view of the above object, the inventors have found that the target di(aminomethyl)-substituted aromatic compound can be produced in a high selectivity by a two-stage hydrogenation comprising a first stage where an aromatic dinitrile is hydrogenated into an intermediate compound (cyano(aminomethyl)-substituted aromatic compound) of the target compound in the presence of a Pd-containing catalyst, and a second stage where the cyano(aminomethyl)-substituted aromatic compound is hydrogenated into the di(aminomethyl)-substituted aromatic compound in the presence of a Ni- and/or Co-containing catalyst. The present invention has been accomplished on the basis of this finding.

[0007] Thus, the present invention provides a method of producing a di(aminomethyl)-substituted aromatic compound represented by the following formula I:



wherein R is a bivalent aromatic group and optionally substituted by a group inert to hydrogenation, by a two-stage hydrogenation comprising a first stage of hydrogenating an aromatic dinitrile represented by the following formula II:



wherein R is the same as defined above, into a cyano(aminomethyl)-substituted aromatic compound represented by the following formula III:



wherein R is the same as defined above, in the presence of a catalyst X containing Pd, and a sec-

ond stage of hydrogenating the cyano(aminomethyl)-substituted aromatic compound of the formula III from the first stage into the di(aminomethyl)-substituted aromatic compound in the presence of a catalyst Y containing Ni and/or Co.

#### DETAILED DESCRIPTION OF THE INVENTION

[0008] The present invention will be described below in more detail.

[0009] The starting aromatic dinitrile is represented by the following formula II:



R is a bivalent aromatic group such as phenylene group, naphthylene group, etc. The positions of two cyano groups on the aromatic group are not particularly limited, for example, may be any one of 0-, m- and p-position if the aromatic group is phenylene group. The aromatic group may be substituted by a group inert to the hydrogenation such as alkyl group, alkoxy group, halogen, amino group, amido group and hydroxyl group. Generally, the reactivity of the aromatic dinitrile in the hydrogenation is largely changed by the substituent group on the aromatic ring. In the method of the present invention, however, the hydrogenation proceeds efficiently even in the presence of such a substituent group. The aromatic dinitrile preferably used in the present invention may be phthalonitrile, isophthalonitrile, terephthalonitrile and 1,5-dicyanonaphthalene.

[0010] Hydrogen to be used in the hydrogenation of the present invention is not required to be specifically purified, and hydrogen of industrial grade is sufficient for the purpose. The hydrogen partial pressure in the reaction system is preferably 2.0 to 20.0 MPa for both the first and second stage hydrogenation. Within the above range, the target diamine is preferably produced in sufficiently high yield and the use of a high-pressure reactor can be avoided to preferably reduce production costs.

[0011] In the first stage hydrogenation, the aromatic dinitrile of the formula II is hydrogenated into the cyano(aminomethyl)-substituted aromatic compound represented by the following formula III:



wherein R is the same as defined above, in the presence of a catalyst X containing Pd.

[0012] The hydrogenation is performed preferably in liquid phase. The solvent to be used is not particularly limited as far as it is an inert organic solvent which is not reduced by hydrogen. Examples thereof include alcohol solvents such as methanol, ethanol and propyl alcohol; hydrocarbon solvents such as m-xylene, mesitylene

and pseudocumene; and ether solvents such as dioxane. The inert organic solvent is preferably used in an amount of 1.0 to 99.0 parts by weight per one part by weight of the aromatic dinitrile. To prevent the formation of by-product, liquid ammonia may be used as the solvent alone or in a mixture with the inert organic solvent mentioned above. The amount of liquid ammonia to be used is preferably 0.5 to 99 parts by weight per one part by weight of the aromatic dinitrile. Within the above range, the lowering of the yield of cyano(aminomethyl)-substituted aromatic compound due to the formation of by-product and the lowering of the space time yield can be preferably prevented. The mixing ratio of the inert organic solvent to liquid ammonia, if used, is preferably 0.01 to 99.0 by weight.

[0013] The first stage hydrogenation may be performed in either batch or continuous manner. The reaction temperature is preferably 20 to 150°C. Within this range, the yield of target compound is preferably increased because the conversion of the starting aromatic dinitrile is high and the by-production of high-boiling compound is prevented.

[0014] The contact time between the starting aromatic dinitrile and the catalyst X is preferably 0.01 to 10.0 h, although depending on the kind of the starting compound, the charged amounts of the starting compound, solvent and hydrogen, the reaction temperature and the reaction pressure.

[0015] The catalyst X used in the first stage for hydrogenating the aromatic dinitrile into the corresponding cyano(aminomethyl)-substituted aromatic compound may be a known catalyst containing Pd. Generally, a catalyst containing Pd supported on  $\text{Al}_2\text{O}_3$ ,  $\text{SiO}_2$ , diatomaceous earth,  $\text{SiO}_2\text{-Al}_2\text{O}_3$  or  $\text{ZrO}_2$ , preferably  $\text{Al}_2\text{O}_3$ , is used as the catalyst X. The supported amount of Pd is preferably 0.05 to 10% by weight based on the catalyst X. The amount of the catalyst X to be used is, in terms of Pd, 0.0001 to 0.1 part by weight per one part by weight of the starting aromatic dinitrile. If the hydrogenation is performed in a fixed bed flow reactor, the flow rate of the starting aromatic dinitrile is preferably 1.0 to 2000  $\text{h}^{-1}$  per unit weight of Pd. By using the catalyst X in an amount within the above range, the hydrogenation proceeds efficiently without increasing the catalyst cost.

[0016] In the second stage hydrogenation, the cyano(aminomethyl)-substituted aromatic compound from the first stage is hydrogenated into the di(aminomethyl)-substituted aromatic compound represented by the following formula I:



wherein R is the same as defined above, in the presence of a catalyst Y containing Ni and/or Co.

[0017] The hydrogenation is performed preferably in liquid phase. The solvent to be used is not particularly

limited as far as it is an inert organic solvent which is not reduced by hydrogen. Examples thereof include alcohol solvents such as methanol, ethanol and propyl alcohol; hydrocarbon solvents such as m-xylene, mesitylene and pseudocumene; and ether solvents such as dioxane. The inert organic solvent is preferably used in an amount of 1.0 to 99.0 parts by weight per one part by weight of the cyano(aminomethyl)-substituted aromatic compound. To prevent the formation of by-product, liquid ammonia may be used as the solvent alone or in a mixture with the inert organic solvent mentioned above. The amount of liquid ammonia to be used is preferably 0.5 to 99 parts by weight per one part by weight of the cyano(aminomethyl)-substituted aromatic compound. Within the above range, the lowering of the yield of di(aminomethyl)-substituted aromatic compound due to the formation of by-product and the lowering of the space time yield can be preferably prevented. The mixing ratio of the inert organic solvent to liquid ammonia, if used, is preferably 0.01 to 99.0 by weight.

**[0018]** The second stage hydrogenation may be performed in either batch or continuous manner. The reaction temperature is preferably 20 to 150°C. Within this range, the yield of target compound is preferably increased because the conversion of the cyano(aminomethyl)-substituted aromatic compound is high and the by-production of high-boiling compound is prevented.

**[0019]** The contact time between the cyano(aminomethyl)-substituted aromatic compound and the catalyst Y is preferably 0.01 to 10.0 h, although depending on the kind of the starting compound, the charged amounts of the starting compound, solvent and hydrogen, the reaction temperature and the reaction pressure.

**[0020]** The catalyst Y used in the first stage for hydrogenating the cyano(aminomethyl)-substituted aromatic compound into the corresponding di(aminomethyl)-substituted aromatic compound may be a known catalyst containing Ni and/or Co. Generally, a catalyst prepared by supporting Ni and/or Co on  $\text{Al}_2\text{O}_3$ ,  $\text{SiO}_2$ , diatomaceous earth,  $\text{SiO}_2\text{-Al}_2\text{O}_3$  or  $\text{ZrO}_2$  by a precipitation method, Raney Ni catalyst or Raney Co catalyst is preferably used as the catalyst Y. The supported amount of the catalytic metal (Ni and/or Co) is preferably 5.0 to 90.0% by weight based on the catalyst Y. The amount of the catalyst Y to be used is, in terms of the catalytic metal, 0.1 to 2.0 parts by weight per one part by weight of the starting cyano(aminomethyl)-substituted aromatic compound. If the hydrogenation is performed in a fixed bed flow reactor, the flow rate of the starting cyano(aminomethyl)-substituted aromatic compound is preferably 0.05 to 5.0  $\text{h}^{-1}$  per unit weight of the catalytic metal. By using the catalyst Y in an amount within the above range, the hydrogenation proceeds efficiently without increasing the catalyst cost.

**[0021]** If the first stage hydrogenation is performed in a reactor other than fixed bed reactor, the reaction prod-

uct solution should be separated from the catalyst X before the resultant cyano(aminomethyl)-substituted aromatic compound is subjected to the second stage hydrogenation. The use of the same solvent in the first stage and the second stage makes the production efficient, because the reaction product solution as separated after the first stage hydrogenation can be directly used in the second stage hydrogenation.

**[0022]** The di(aminomethyl)-substituted aromatic compound produced in the second stage hydrogenation is recovered after separated from the solvent and the catalyst in known manners. For example, after separating the reaction system into the vapor component and the liquid component, the target compound is recovered from the liquid component by distillation.

**[0023]** In the known production methods of the di(aminomethyl)-substituted aromatic compound, high-boiling by-products which are formed during the reaction adhere to the catalyst thereby to gradually increase the pressure difference. Therefore, the catalyst should be regenerated by decomposing the high-boiling by-products by hydrogenation. Thus, the catalyst life is shortened in the known production methods. In contrast, the catalyst life is prolonged in the production method of the present invention, because the formation of by-products is minimized.

**[0024]** The present invention will be explained in more detail by reference to the following example which should not be construed to limit the scope of the present invention.

#### Preparation of Ni Catalyst

**[0025]** An aqueous solution of mixed metal salts was prepared by dissolving 305.0 g of nickel nitrate hexahydrate ( $\text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ ) and 13.6 g of cobalt nitrate hexahydrate ( $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ ) in 840 g of pure water at 40°C. Separately, 190.6 g of ammonium hydrogencarbonate ( $\text{NH}_4\text{HCO}_3$ ) was dissolved in 2.4 kg of pure water under sufficient stirring while raising the temperature to 40°C. To the aqueous solution of ammonium hydrogencarbonate, the aqueous solution of mixed metal salts kept at 40°C was added under sufficient stirring to prepare a precipitation slurry of nickel carbonate. The slurry was heated to 80°C and kept there for 30 min. Then, the slurry was cooled to 40°C and kept there. Separately, 118.4 g of zirconium nitrate aqueous solution (Zr content: 25% by weight in terms of  $\text{ZrO}_2$ ) was mixed with 300 g of pure water and kept at 40°C. Further, 42.8 g of ammonium hydrogencarbonate ( $\text{NH}_4\text{HCO}_3$ ) was dissolved in 530 g of pure water and kept at 40°C. The aqueous solution of zirconium nitrate and the aqueous solution of ammonium hydrogencarbonate were simultaneously added to the precipitation slurry of nickel carbonate to precipitate zirconium carbonate. The precipitation slurry thus prepared was stirred for 30 min while being kept at 40°C. Then, the precipitation slurry was filtered. The separated precipitation was washed, dried

at 110°C overnight and calcined at 380°C for 18 h in air. The resultant calcined powder was mixed with 3% by weight of graphite and made into 3.0 mm  $\phi$   $\times$  2.5 mm tablets, which were then reduced at 400°C in hydrogen stream to prepare catalyst A. The supported amount of Ni in the catalyst A was 65% by weight.

#### EXAMPLE 1 Hydrogenation of Isophthalonitrile

[0026] Into a 100-ml autoclave, were charged 3.2 g of isophthalonitrile, 10.4 g of mesitylene, 10.0 g of liquid ammonia and 2.0 g of Pd-alumina pellets (manufactured by N.E. Chemcat Corporation; Pd content = 5% by weight), and the inner pressure was raised to 4.9 MPa by hydrogen gas. Then, the autoclave was shaken at 50°C until the change of pressure was no longer appreciated. The analysis on the reaction product solution showed that the conversion of isophthalonitrile was 95.7 mol%, the yield of 3-cyanobenzylamine was 87.3 mol% and the yield of m-xylylenediamine was 7.7 mol%. The reaction solution separated from the catalyst was charged into a 100-ml autoclave together with 10.0 g of liquid ammonia and 2.0 g of Ni-diatomaceous earth pellets (manufactured by Nikki Chemical Co., Ltd.; Ni supported amount = 46% by weight). The inner pressure was raised to 4.9 MPa by hydrogen gas. Then, the autoclave was shaken at 50°C until the change of pressure was no longer appreciated. The analysis on the reaction product solution showed that the conversion of isophthalonitrile was 100 mol%, the yield of 3-cyanobenzylamine was 0.2 mol% and the yield of m-xylylenediamine was 89.4 mol%.

#### EXAMPLE 2 Hydrogenation of Terephthalonitrile

[0027] Into a 100-ml autoclave, were charged 3.2 g of terephthalonitrile, 10.4 g of mesitylene, 10.0 g of liquid ammonia and 2.0 g of Pd-alumina pellets (manufactured by N.E. Chemcat Corporation; Pd content = 5% by weight), and the inner pressure was raised to 4.9 MPa by hydrogen gas. Then, the autoclave was shaken at 50°C until the change of pressure was no longer appreciated. The analysis on the reaction product solution showed that the conversion of terephthalonitrile was 94.8 mol%, the yield of 4-cyanobenzylamine was 88.8 mol% and the yield of p-xylylenediamine was 5.8 mol%. The reaction solution separated from the catalyst was charged into a 100-ml autoclave together with 10.0 g of liquid ammonia and 2.0 g of Ni-diatomaceous earth pellets (manufactured by Nikki Chemical Co., Ltd.; Ni supported amount = 46% by weight). The inner pressure was raised to 4.9 MPa by hydrogen gas. Then, the autoclave was shaken at 50°C until the change of pressure was no longer appreciated. The analysis on the reaction product solution showed that the conversion of terephthalonitrile was 100 mol%, the yield of 4-cyanobenzylamine was 0.5 mol% and the yield of p-xylylenediamine was 87.7 mol%.

#### EXAMPLE 3 Hydrogenation of 1,5-Dicyanonaphthalene

[0028] Into a 100-ml autoclave, were charged 3.2 g of 1,5-dicyanonaphthalene, 10.4 g of mesitylene, 10.0 g of liquid ammonia and 2.0 g of Pd-alumina pellets (manufactured by N.E. Chemcat Corporation; Pd content = 5% by weight), and the inner pressure was raised to 4.9 MPa by hydrogen gas. Then, the autoclave was shaken at 50°C until the change of pressure was no longer appreciated. The analysis on the reaction product solution showed that the conversion of 1,5-dicyanonaphthalene was 92.6 mol%, the yield of 1-aminomethyl-5-cyanonaphthalene was 85.4 mol% and the yield of 1,5-diaminomethylnaphthalene was 4.0 mol%. The reaction solution separated from the catalyst was charged into a 100-ml autoclave together with 10.0 g of liquid ammonia and 2.0 g of Ni-diatomaceous earth pellets (manufactured by Nikki Chemical Co., Ltd.; Ni supported amount = 46% by weight). The inner pressure was raised to 4.9 MPa by hydrogen gas. Then, the autoclave was shaken at 50°C until the change of pressure was no longer appreciated. The analysis on the reaction product solution showed that the conversion of 1,5-dicyanonaphthalene was 100 mol%, the yield of 1-aminomethyl-5-cyanonaphthalene was 0.0 mol% and the yield of 1,5-diaminomethylnaphthalene was 88.0 mol%.

#### EXAMPLE 4 Hydrogenation of Isophthalonitrile

[0029] Into a 100-ml autoclave, were charged 3.2 g of isophthalonitrile, 10.4 g of mesitylene, 10.0 g of liquid ammonia and 2.0 g of Pd-alumina pellets (manufactured by N.E. Chemcat Corporation; Pd content = 5% by weight), and the inner pressure was raised to 4.9 MPa by hydrogen gas. Then, the autoclave was shaken at 50°C until the change of pressure was no longer appreciated. The analysis on the reaction product solution showed that the conversion of isophthalonitrile was 95.7 mol%, the yield of 3-cyanobenzylamine was 87.3 mol% and the yield of m-xylylenediamine was 7.7 mol%. The reaction solution separated from the catalyst was charged into a 100-ml autoclave together with 10.0 g of liquid ammonia and 2.0 g of the catalyst A. The inner pressure was raised to 4.9 MPa by hydrogen gas. Then, the autoclave was shaken at 50°C until the change of pressure was no longer appreciated. The analysis on the reaction product solution showed that the conversion of isophthalonitrile was 100 mol%, the yield of 3-cyanobenzylamine was 0.0 mol% and the yield of m-xylylenediamine was 91.1 mol%.

#### EXAMPLE 5 Hydrogenation of Terephthalonitrile

[0030] Into a 100-ml autoclave, were charged 3.2 g of terephthalonitrile, 10.4 g of mesitylene, 10.0 g of liquid ammonia and 2.0 g of Pd-alumina pellets (manufactured by N.E. Chemcat Corporation; Pd content = 5% by weight), and the inner pressure was raised to 4.9 MPa

by hydrogen gas. Then, the autoclave was shaken at 50°C until the change of pressure was no longer appreciated. The analysis on the reaction product solution showed that the conversion of terephthalonitrile was 94.8 mol%, the yield of 4-cyanobenzylamine was 88.8 mol% and the yield of p-xylylenediamine was 5.8 mol%. The reaction solution separated from the catalyst was charged into a 100-ml autoclave together with 10.0 g of liquid ammonia and 2.0 g of the catalyst A. The inner pressure was raised to 4.9 MPa by hydrogen gas. Then, the autoclave was shaken at 50°C until the change of pressure was no longer appreciated. The analysis on the reaction product solution showed that the conversion of terephthalonitrile was 100 mol%, the yield of 4-cyanobenzylamine was 0.2 mol% and the yield of p-xylylenediamine was 92.1 mol%.

#### EXAMPLE 6 Hydrogenation of 1,5-Dicyanonaphthalene

[0031] Into a 100-ml autoclave, were charged 3.2 g of 1,5-dicyanonaphthalene, 10.4 g of mesitylene, 10.0 g of liquid ammonia and 2.0 g of Pd-alumina pellets (manufactured by N.E. Chemcat Corporation; Pd content = 5% by weight), and the inner pressure was raised to 4.9 MPa by hydrogen gas. Then, the autoclave was shaken at 50°C until the change of pressure was no longer appreciated. The analysis on the reaction product solution showed that the conversion of 1,5-dicyanonaphthalene was 92.6 mol%, the yield of 1-aminomethyl-5-cyanonaphthalene was 85.4 mol% and the yield of 1,5-diaminomethylnaphthalene was 4.0 mol%. The reaction solution separated from the catalyst was charged into a 100-ml autoclave together with 10.0 g of liquid ammonia and 2.0 g of the catalyst A. The inner pressure was raised to 4.9 MPa by hydrogen gas. Then, the autoclave was shaken at 50°C until the change of pressure was no longer appreciated. The analysis on the reaction product solution showed that the conversion of 1,5-dicyanonaphthalene was 100 mol%, the yield of 1-aminomethyl-5-cyanonaphthalene was 1.5 mol% and the yield of 1,5-diaminomethylnaphthalene was 87.1 mol%.

#### COMPARATIVE EXAMPLE 1 Hydrogenation of Isophthalonitrile

[0032] Into a 100-ml autoclave, were charged 3.2 g of isophthalonitrile, 10.4 g of mesitylene, 10.0 g of liquid ammonia and 2.0 g of the catalyst A, and the inner pressure was raised to 4.9 MPa by hydrogen gas. Then, the autoclave was shaken at 50°C until the change of pressure was no longer appreciated. The analysis on the reaction product solution showed that the conversion of isophthalonitrile was 95.5 mol% and the yield of m-xylylenediamine was 49.4 mol%.

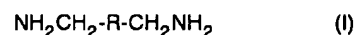
#### COMPARATIVE EXAMPLE 2 Hydrogenation of Terephthalonitrile

[0033] Into a 100-ml autoclave, were charged 3.2 g of terephthalonitrile, 10.4 g of mesitylene, 10.0 g of liquid ammonia and 2.0 g of the catalyst A, and the inner pressure was raised to 4.9 MPa by hydrogen gas. Then, the autoclave was shaken at 50°C until the change of pressure was no longer appreciated. The analysis on the reaction product solution showed that the conversion of terephthalonitrile was 94.4 mol% and the yield of p-xylylenediamine was 35.6 mol%.

[0034] According to the present invention, the di(aminomethyl)-substituted aromatic compound is produced in a high selectivity and a sufficiently high yield by the two-stage hydrogenation comprising the first stage where the aromatic dinitrile is hydrogenated into an intermediate compound (cyano(aminomethyl)-substituted aromatic compound) of the target compound in the presence of a Pd-containing catalyst, and a second stage where the cyano(aminomethyl)-substituted aromatic compound is hydrogenated into the di(aminomethyl)-substituted aromatic compound in the presence of a Ni- and/or Co-containing catalyst. In addition, the catalyst life is prolonged in the production method of the present invention, because the formation of by-products is minimized. Therefore, by the production method of the present invention, the di(aminomethyl)-substituted aromatic compound is produced from the aromatic dinitrile at low production costs.

#### Claims

1. A method of producing a di(aminomethyl)-substituted aromatic compound represented by the following formula I:



wherein R is a bivalent aromatic group and optionally substituted by a group inert to hydrogenation, by a two-stage hydrogenation comprising:

a first stage of hydrogenating an aromatic dinitrile represented by the following formula II:



wherein R is the same as defined above, into a cyano(aminomethyl)-substituted aromatic compound represented by the following formula III:



wherein R is the same as defined above, in the presence of a catalyst X containing Pd; and  
a second stage of hydrogenating the cyano (aminomethyl)-substituted aromatic compound from the first stage into the di(aminomethyl)-substituted aromatic compound in the presence of a catalyst Y containing Ni and/or Co.

2. The method according to Claim 1, wherein the catalyst X comprises 0.05 to 10% by weight of Pd carried on a support.
3. The method according to Claim 2, wherein the support is alumina.
4. The method according to Claim 1, wherein the catalyst Y is prepared by a precipitation method to carry Ni and/or Co on a support.
5. The method according to Claim 1, wherein the catalyst Y is a Raney nickel catalyst of a Raney cobalt catalyst.



European Patent  
Office

## EUROPEAN SEARCH REPORT

Application Number  
EP 04 00 2100

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.7)
D,A	DATABASE CHEMABS [Online] CHEMICAL ABSTRACTS SERVICE, COLUMBUS, OHIO, US; ONDA, YUJI ET AL: "Cyanobenzylamine" XP002276586 retrieved from STN Database accession no. 1975:4040 * abstract * & JP 510 244 94B B4 (MITSUBISHI GAS CHEMICAL CO., INC.) 24 July 1976 (1976-07-24) -----	1-5	C07C253/30 C07C209/48 C07C211/27 C07C211/30
D,A	DATABASE CHEMABS [Online] CHEMICAL ABSTRACTS SERVICE, COLUMBUS, OHIO, US; MORIKAWA, KOHEI ET AL: "Preparation of aromatic cyano methylamines with selective hydrogenation catalysts" XP002276587 retrieved from STN Database accession no. 1997:264560 * abstract * & JP 09 040630 A2 (SHOWA DENKO KK, JAPAN) 10 February 1997 (1997-02-10) -----	1-5	TECHNICAL FIELDS SEARCHED (Int.Cl.7) C07C
D,A	DATABASE CHEMABS [Online] CHEMICAL ABSTRACTS SERVICE, COLUMBUS, OHIO, US; MIURA, KANEO ET AL: "Preparation of aromatic cyanomethylamines from aromatic dinitriles" XP002276588 retrieved from STN Database accession no. 1998:498623 * abstract * & JP 10 204048 A2 (SHOWA DENKO K. K., JAPAN) 4 August 1998 (1998-08-04) ----- -/--	1-5	
The present search report has been drawn up for all claims			
Place of search The Hague		Date of completion of the search 14 April 2004	Examiner Beslier, L
CATEGORY OF CITED DOCUMENTS X: particularly relevant if taken alone Y: particularly relevant if combined with another document of the same category A: technological background O: non-written disclosure P: intermediate document		T: theory or principle underlying the invention E: earlier patent document, but published on, or after the filing date D: document cited in the application L: document cited for other reasons & : member of the same patent family, corresponding document	

EPO FORM 1503 03.92 (P04.001)





European Patent  
Office

## EUROPEAN SEARCH REPORT

Application Number  
EP 04 00 2100

DOCUMENTS CONSIDERED TO BE RELEVANT				
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.7)	
D,A	WO 96/18603 A (RHONE-POULENC FIBER AND RESINE INTERMEDIATES) 20 June 1996 (1996-06-20) * the whole document *	1-5		
D,A	----- DATABASE CHEMABS [Online] CHEMICAL ABSTRACTS SERVICE, COLUMBUS, OHIO, US; KAWAKATSU, MICHIO ET AL: "Xylylenediamine by hydrogenation of phthalonitrile" XP002276589 retrieved from STN Database accession no. 1963:461849 * abstract * & JP 38 008719 19 (TOHO RAYON CO., LTD.) 21 October 1960 (1960-10-21)	1-5		
D,A	----- DATABASE CHEMABS [Online] CHEMICAL ABSTRACTS SERVICE, COLUMBUS, OHIO, US; AYAKI, KAZUO ET AL: "Nickel system reducing catalyst" XP002276590 retrieved from STN Database accession no. 1974:19925 * abstract * & JP 480 225 93B B4 (SHOWA DENKO K. K.) 6 July 1973 (1973-07-06)	1-5		TECHNICAL FIELDS SEARCHED (Int.Cl.7)
A	----- US 6 476 267 B1 (TAKAMASA FUCHIGAMI) 5 November 2002 (2002-11-05) * examples 21,22 *	1-5		
A	----- DD 77 983 A (DR. HANS BALTZ) 5 December 1970 (1970-12-05) * the whole document *	1-5		
		-/--		
The present search report has been drawn up for all claims				
Place of search The Hague		Date of completion of the search 14 April 2004	Examiner Beslier, L	
CATEGORY OF CITED DOCUMENTS X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons & : member of the same patent family, corresponding document				

EPO FORM 1500 (03.02 (1994.01))



European Patent  
Office

# EUROPEAN SEARCH REPORT

Application Number  
EP 04 00 2100

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.7)
A	DATABASE CHEMABS [Online] CHEMICAL ABSTRACTS SERVICE, COLUMBUS, OHIO, US; OKAWA, TAKASHI: "Process for manufacturing aromatic cyanoaminomethyl compounds" XP002276591 retrieved from STN Database accession no. 2002:547258 * abstract * & JP 2002 205980 A2 (MITSUBISHI GAS CHEMICAL CO., LTD., JAPAN) 23 July 2002 (2002-07-23) -----	1-5	
A	US 3 647 054 A (YUKIO TSUBOI) 7 March 1972 (1972-03-07) * the whole document *	1-5	
A	EP 0 908 447 A (SHOWA DENKO K.K.) 14 April 1999 (1999-04-14) * the whole document *	1-5	
P,A	DATABASE CHEMABS [Online] CHEMICAL ABSTRACTS SERVICE, COLUMBUS, OHIO, US; AMAKAWA, KAZUHIKO ET AL: "Process for manufacturing aromatic dimethylamine" XP002276592 retrieved from STN Database accession no. 2003:906050 * abstract * & JP 2003 327563 A2 (MITSUBISHI GAS CHEMICAL CO., LTD., JAPAN) 19 November 2003 (2003-11-19) -----	1-5	
The present search report has been drawn up for all claims			TECHNICAL FIELDS SEARCHED (Int.Cl.7)
Place of search The Hague		Date of completion of the search 14 April 2004	Examiner Beslier, L
<p>CATEGORY OF CITED DOCUMENTS</p> <p>X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document</p> <p>T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons &amp; : member of the same patent family, corresponding document</p>			

EPO FORM 1503 03.02 (P4001)

**ANNEX TO THE EUROPEAN SEARCH REPORT  
ON EUROPEAN PATENT APPLICATION NO.**

EP 04 00 2100

This annex lists the patent family members relating to the patent documents cited in the above-mentioned European search report.  
The members are as contained in the European Patent Office EDP file on  
The European Patent Office is in no way liable for these particulars which are merely given for the purpose of information.

14-04-2004

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
JP 51024494B B4	15-08-1974	JP 49085041 A	15-08-1974
		JP 51024494 B	24-07-1976
JP 9040630 A2	10-02-1997	JP 2937083 B2	23-08-1999
		JP 9040630 A	10-02-1997
JP 10204048 A2	04-08-1998	JP 10204048 A	04-08-1998
WO 9618603 A	20-06-1996	FR 2728259 A1	21-06-1996
		BR 9510011 A	21-10-1997
		CA 2208249 A1	20-06-1996
		CN 1169717 A ,B	07-01-1998
		DE 69523620 D1	06-12-2001
		DE 69523620 T2	25-04-2002
		EP 0797568 A1	01-10-1997
		WO 9618603 A1	20-06-1996
		JP 2938975 B2	25-08-1999
		JP 10502671 T	10-03-1998
		RU 2167854 C2	27-05-2001
		US 5981790 A	09-11-1999
JP 38008719 19		NONE	
JP 48022593B B4	06-07-1973	JP 48022593 B	06-07-1973
US 6476267 B1	05-11-2002	WO 0046179 A1	10-08-2000
DD 77983 A		NONE	
JP 2002205980 A2	23-07-2002	JP 2002205980 A	23-07-2002
US 3647054 A	07-03-1972	NONE	
EP 0908447 A	14-04-1999	WO 9833767 A1	06-08-1998
		WO 9833766 A1	06-08-1998
		AU 5779998 A	25-08-1998
		EP 0908447 A1	14-04-1999
		IL 126443 A	17-09-2003
		NO 984622 A	23-11-1998
		US 6114277 A	05-09-2000
		AU 1558697 A	25-08-1998
		CA 2250770 A1	06-08-1998
		CN 1216038 A	05-05-1999
		CN 1216038 T	05-05-1999
		CZ 9803522 A3	16-06-1999
		KR 2000044033 A	15-07-2000

EPO FORM P0459

For more details about this annex : see Official Journal of the European Patent Office, No. 12/82

